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https://www.youtube.com/watch?v=w0RfMRDy34w Assign oxidation numbers for all of the elements in the following compounds:		
	Fe_2O_3	
CH_4	NaHCO ₃	



QUESTION In which of the following does nitrogen have an oxidation state of +4? A. HNO₃ B. NO₂ C. NH₄CI D. NaNO₂



B) NO₂

Oxygen has an oxidation state of -2 when part of a compound.





























Answer

TRUE (A) / FALSE (B)

Copper is oxidized in the reaction and silver is reduced.

 $Cu(s) - 2e^- \rightarrow Cu^{2+}(aq)$

Question

 $\begin{array}{ccc} 0 & +1 & 0 & +2 \\ \text{Cu(s)} + 2 \text{ AgNO}_3(\text{aq}) & \rightarrow 2 \text{Ag(s)} + \text{Cu(NO}_3)_2(\text{aq}) \end{array}$

TRUE (A) / FALSE (B)

Copper looses 2 electrons in the reaction and each silver atom gains 1 electron.



