





















Question $Gas \Rightarrow Liquid Equilibrium$ What happens when a container with neon gas and neon liquid is made smaller, while keeping the temperature constant?



- The system will shift to the right producing more liquid. А.
- The system will shift to the left producing more gas. R
- This system would not С. affected.











Question $Gas \Rightarrow Liquid Equilibrium$

What happens when a container with neon gas and neon liquid is cooled, while keeping the pressure constant?

- A. The system will shift to the right producing more liquid.
 B. The system will shift to the left producing more gas.
- left producing more gas. c. This system would not be
- c. This system would not b affected.





























QUESTION

If a 10.0 L sample of a gas at 25 $^{\rm o}{\rm C}$ suddenly had its volume doubled, without changing its temperature what would happen to its pressure? What could be done to keep the pressure constant without changing the temperature?

- A) The pressure would double; nothing else could be done to prevent this.
- B) The pressure would double; the moles of gas could be doubled.C) The pressure would decrease by a factor of two; the moles of gas could be halved.
- D) The pressure would decrease by a factor of two; the moles could be doubled.

Gases & Airbags Use of Chemical Reactions and Physical Properties



http://chemconnections.org/public_html/general/movies/airbags.MOV