Name: _____

Section: _____

Report Form – Name, Ions, and Formula Activity

Name	Separated Ions	Formula
Sodium chloride	1 Na^+ + 1 Cl^-	NaCl
Calcium chloride	1 Ca^{2+} + 2 Cl^-	$CaCl_2$
Lithium carbonate		
Barium hydroxide		
	__ K^+ + __ SO_4^{2-}	
	__NH_4^+ + CO_3^{2-}	
		$FeBr_2$
		$Fe_2(SO_4)_3$
Copper(II) nitrate		
Tin(IV) fluoride		
	__Al^{3+} + __SO_3^{2-}	
		$Ca(NO_2)_2$
		$PbCl_4$
	__Fe^{2+} + __PO_4^{3-}	
		$HgBr_2$
Calcium acetate		
Cobalt(III) sulfate		

Show the **formula** of the following combination of ions.

	Na^+	Ba^{2+}	Cu^+	Fe^{3+}	Al^{3+}	NH_4^+
Cl^-						
NO_3^-						
CO_3^{2-}						
PO_4^{3-}						
SO_3^{2-}						
$C_2H_3O_2^-$						

Report Form – Naming Activity

Write the name for the following combination of ions. (Refer to the previous page for formulas.)

	Na^+	Ba^{2+}	Cu^+	Fe^{3+}	Al^{3+}	NH_4^+
Cl^-						
NO_3^-						
CO_3^{2-}						
PO_4^{3-}						
SO_3^{2-}						
$C_2H_3O_2^-$						

Name: _____

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Report Form – Nomenclature Activity

Write the name for each compound.

Salts and Hydroxides

Na_2S	_____	$HgCl_2$	_____
$AlCl_3$	_____	$CoCl_3$	_____
BaF_2	_____	$Cu(OH)_2$	_____
$Ca_3(PO_4)_2$	_____	$CuOH$	_____
CaO	_____	Fe_2S_3	_____
NH_4Cl	_____	$Zn(C_2H_3O_2)_2$	_____
KOH	_____	PbI_4	_____
Na_2CO_3	_____	HgO	_____
K_2SO_4	_____	Cu_2O	_____
$Ca(NO_2)_2$	_____	$Pb(NO_2)_2$	_____
$Mg(C_2H_3O_2)_2$	_____	$Pb(SO_4)_2$	_____
$Al_2(CO_3)_3$	_____	$KHCO_3$	_____
NH_4I	_____	Na_3PO_4	_____
$BaSO_4$	_____	K_3PO_4	_____
$NaC_2H_3O_2$	_____	$FeCl_3$	_____
KNO_3	_____	$FeCl_2$	_____
$(NH_4)_3PO_4$	_____	$FeSO_4$	_____
Al_2O_3	_____	$PbCl_2$	_____
K_2CO_3	_____	$PbCO_3$	_____
$Mg(NO_3)_2$	_____	$Cu_3(PO_4)_2$	_____

Acids (all aqueous)

HNO_3	_____	H_2SO_4	_____
H_2CrO_4	_____	H_2CO_3	_____
H_3PO_4	_____	H_2S	_____
HCl	_____	$HC_2H_3O_2$	_____
$HClO_3$	_____	HNO_2	_____

Covalent Compounds

B_2O_3	_____	N_2Cl_4	_____
SO_3	_____	N_2O_5	_____
NCl_3	_____	ClF_5	_____
P_2O_5	_____	CCl_4	_____

Write the correct formula for the following compounds:

sodium bromide	_____	hydrosulfuric acid	_____
calcium chloride	_____	iron(II) hydroxide	_____
calcium sulfide	_____	lead(II) chloride	_____
copper(I) carbonate	_____	calcium hydrogen carbonate	_____
sulfuric acid	_____	potassium phosphate	_____
diphosphorus pentoxide	_____	sulfur trioxide	_____
chloric acid	_____	phosphoric acid	_____
ammonium sulfate	_____	bromine pentafluoride	_____
hydrochloric acid	_____	dinitrogen tetroxide	_____
mercury(II) acetate	_____	carbon monoxide	_____
lead (II) nitrite	_____	aluminum carbonate	_____
ammonium sulfite	_____	ammonium carbonate	_____
sodium sulfate	_____	cobalt (II) phosphate	_____
nitrous acid	_____	barium phosphate	_____
nitric acid	_____	mercury(II) bromide	_____
iron(III) sulfate	_____	sodium sulfite	_____
barium fluoride	_____	magnesium phosphate	_____
copper(I) sulfite	_____	cobalt(II) nitrate	_____
ammonium nitrate	_____	iron(III) oxide	_____
iron(III) hydroxide	_____	dinitrogen pentoxide	_____
mercury(II) nitrate	_____	hydrobromic acid	_____
sulfurous acid	_____	magnesium sulfate	_____
lead(IV) phosphate	_____	nickel(II) hydrogen carbonate	_____
sulfur hexafluoride	_____	ammonium phosphate	_____